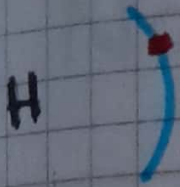


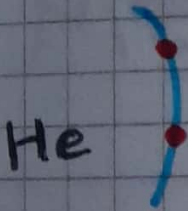
PROPIEDADES PERIÓDICAS DE LOS ELEMENTOS

1. HIDRÓGENO:



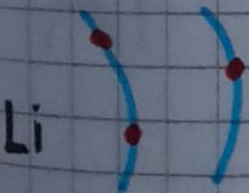
Nº: 1

2. HELIO



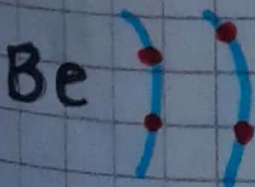
Nº: 1

3. LITIO:



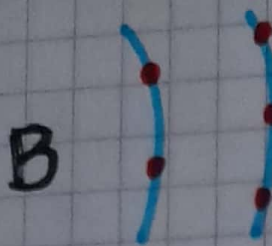
Nº: 2

4. BERILIO



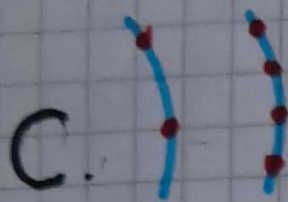
Nº: 2

• 5: BORO



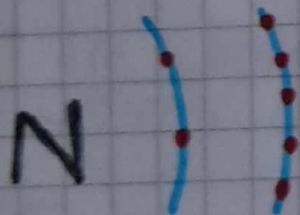
Nº: 2.

• 6: CARBON



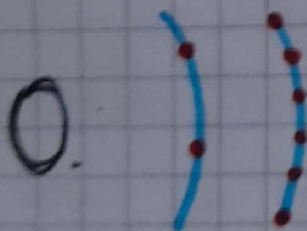
Nº: 2.

• 7: NITROGENO



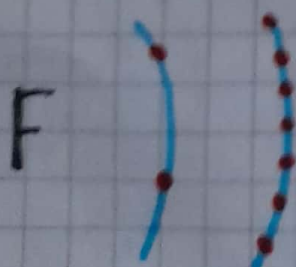
Nº: 2.

• 8: OXIGENO



Nº: 2.

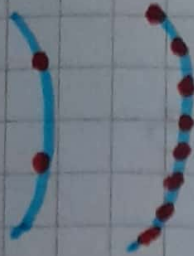
• 9: FLUOR



Nº: 2

10: NEON

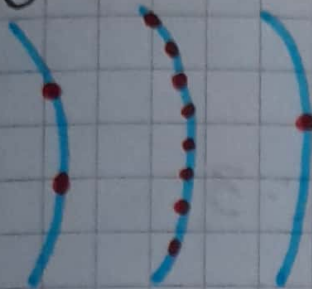
Ne:



N: 2.

11: SODIO

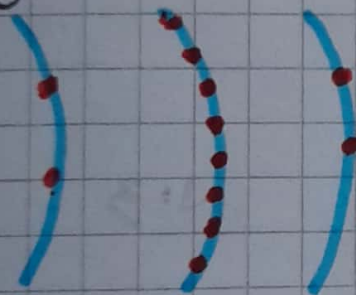
Na:



N: 3

12: MAGNESIO

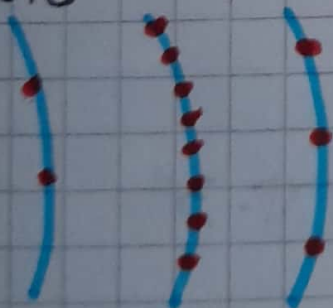
Mg



N: 3

13: ALUMINIO

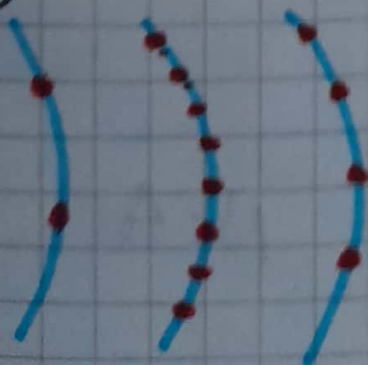
Al



N: 3

14: SILICIO

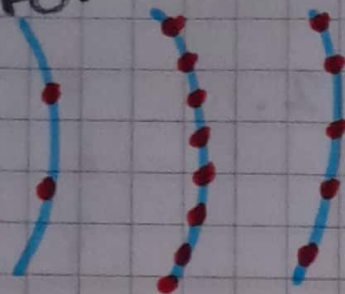
Si



N: 3

15: FOSFORO:

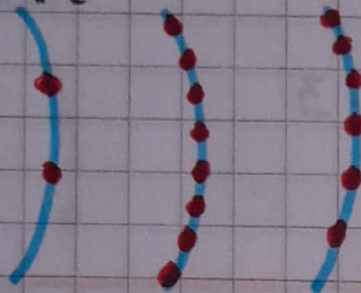
P.



Nº: 3

16: AZUFRE

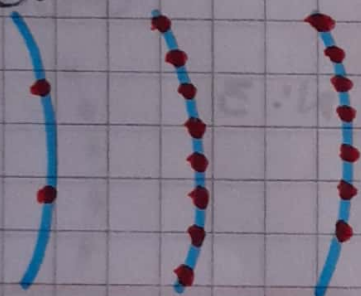
S



Nº: 3

17: CLORO:

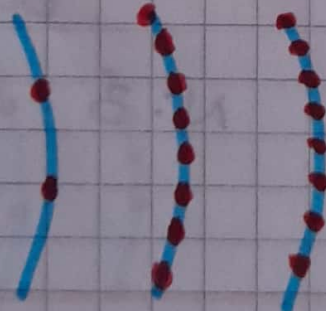
Cl



Nº: 3

18: ARGON:

Ar



Nº: 3

19: POTASIO:

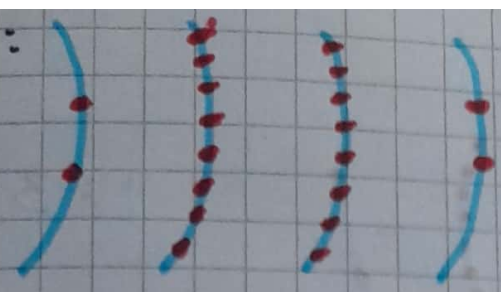
K



Nº: 4

20: Calcio:

Ca:



N: 4

21: ESCANDIO

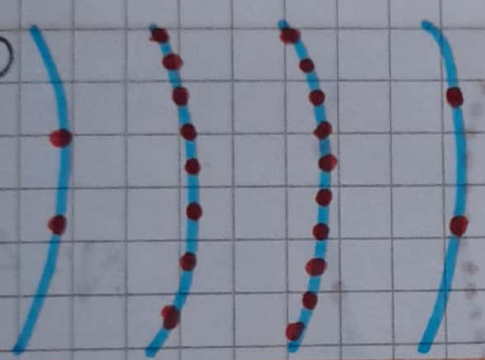
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N: 4.

22: TITANIO

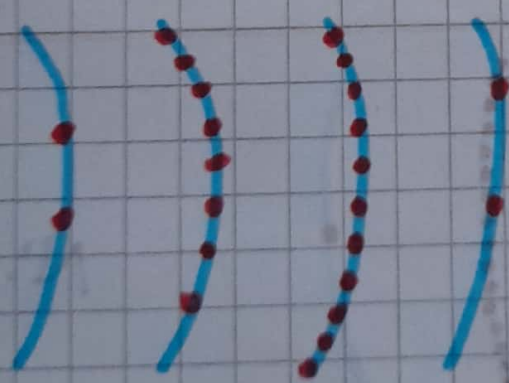
Ti



N: 4

23: VANADIO:

V.



N: 4

24: Cromo

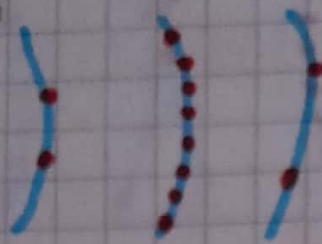
Cr.



N: 4

• 25: MANGANESE

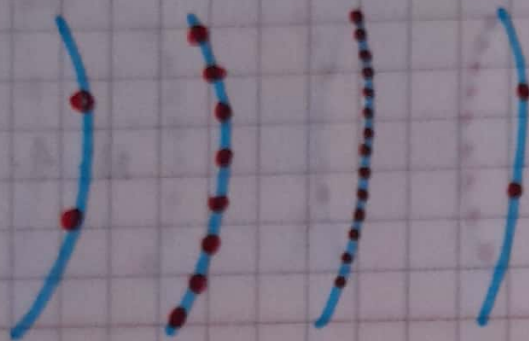
Mn



Nº 3

• 26: Hierro

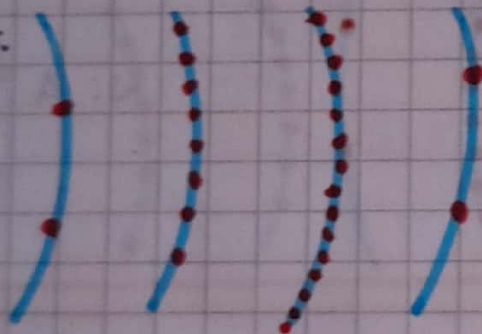
Fe



Nº 4

• 27: COBALTO:

Co



Nº 4

• 28: NIQUEL

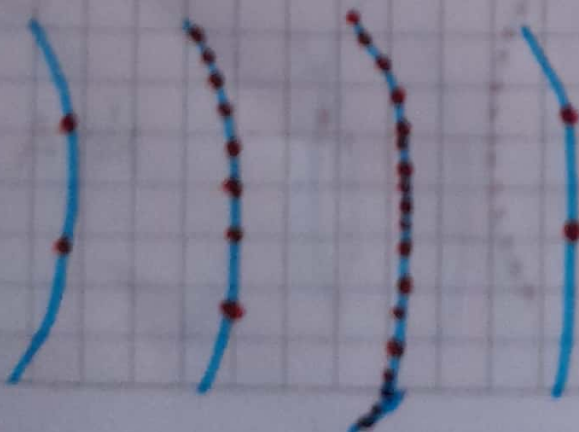
Ni



Nº 4

• 29: COBRE

Cu



Nº 4

30: 2inc

Z_n



N: A



INSTITUCIÓN EDUCATIVA MIGUEL DE CERVANTES SAAVEDRA

CATIONES Y ANIONES						
Ion	Tipo	No. Atómico (Z)	Masa atómica (A)	Protones (p ⁺)	Electrones (e ⁻)	Neutrones n ^o
Na ⁺	cación	11	23	11	10	12
S ⁻²	ANION	16	32	16	18	16
Mg ⁺²	CATION	12	24	12	10	12
Ga ⁺³	CATION	31	70	31	28	39
Se ⁻²	ANION	34	79	34	36	45
Al ⁺³	CATION	13	27	13	10	14
Br ⁻	anión	35	80	35	36	45
Fr ⁺	CATION	87	223	87	86	136
Cl ⁻	ANION	17	35	17	17	18
F ⁻	ANION	9	19	9	10	10
Sr ⁺²	CATION	38	85	38	38	47
B ⁺³	CATION	5	11	5	2	6